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Chemistry 101 ANSWER KEY 1 REVIEW QUESTIONS Chapter 9 1. Draw Lewis structures for each of the following structures and assign formal charges to each atom: a) SF<sub>2</sub> 20 electrons 0 0 0 b) NH<sub>2</sub> OH (N and O are bonded to one another) 14 electrons 0 0 0 0 0 c) PO<sub>4</sub><sup>3-</sup>

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Answer Key Chapter 9 - Chemistry 2e | OpenStax. 1. The cutting edge of a knife that has been sharpened has a smaller surface area than a dull knife. Since pressure is force per unit area, a sharp knife will exert a higher pressure with the same amount of force and cut through material more effectively. 3.

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show a chemical reaction, using symbols a.... a process in which one or more substances are changed into dif.... number in front of a chemical formula in an equation that indi....

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one or more substances are changed into dif.... number in front of a chemical formula in an equation that indi....

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8 Chapter 1 Review Solutions 9. It does not. A significant figure is one that has been measured. It doesn't have to be mathematically important. For example, the zero in 1.0 is significant, but it is not mathematically important, as 1.0 and 1 have the same mathematical value. 10. There is always uncertainty in the last significant figure.

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Chemistry (4th Edition) Burdge, Julia Publisher McGraw-Hill Publishing Company ISBN 978-0-07802-152-7

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Chapter 9: Standard Review Worksheet 1. Answers will vary. An example is included below:  $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$  This describes the decomposition reaction of hydrogen peroxide.

Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to produce two molecules of liquid water and one molecule of oxygen gas.

## Chapter 9: Standard Review Worksheet

CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW SHORT

ANSWER Answer the following questions in the space provided.

1. Given the following equation:  $\text{C}_3\text{H}_8(\text{g}) + x\text{O}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g})$

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+ 2H<sub>2</sub>O(g) 4 a. What is the value of the coefficient x in this equation? 40.07 g/mol b. What is the molar mass of C<sub>3</sub>H<sub>4</sub>? 2 mol O<sub>2</sub>:1 mol H<sub>2</sub>O c. What is the mole ratio of O<sub>2</sub> to H

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9. Elements can gain the stable atomic structure of a noble gas



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by either gaining or losing electrons. Two examples are sodium, which, when it loses an electron, has the same atomic structure as neon, a noble gas; and chlorine, which when it gains an electron, has the same atomic structure as argon, another noble gas.

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