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Chapter 21: ELECTROCHEMISTRY TYING IT ALL TOGETHER-RT In $K = G = -nFE_0$ CH 17-20 = CH15 = CH 21. 2 Important errata to fix. $F = 9.649 \times 10^4$ C/mole in problem on current measurement. Beginning of section on anodes and cathodes should read: "Oh, and by

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Chapter 21 Notes Electrochemical Cells Page 1 of 4 21.1 Electrochemical Cells Electrochemical Processes o A Spontaneous Redox Reaction KEY = For any two metals in an activity series, the more active metal is more readily oxidized than the less active metal o Redox Reactions and Electrochemistry

Chapter 21 Notes Electrochemical Cells - Mr. Miller

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1. In an electrolytic cell the electrode at which the electrons

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enter the solution is called the ____ ; the chemical change that occurs at this electrode is called ____ . (a) anode, oxidation (b) anode, reduction (c) cathode, oxidation (d) cathode, reduction (e) cannot tell unless we know the species being oxidized and reduced.

Sample Questions - Chapter 21

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Chapter 21: ELECTROCHEMISTRY TYING IT ALL TOGETHER On the basis of these marks the correct answer will be. (a) In dilute sulphuric acid solution, hydrogen will be reduced at ... Electrochemistry - Answers Answer: (a) tin electrode is the cathode; cathode is the site of reduction

